

Article

Unraveling Toluene Conversion during the Liquid Phase Hydrogenation of Cyclohexene (in Toluene) with Rh Hybrid Catalysts

Mónica Rufete-Beneite and M. Carmen Román-Martínez *

Department of Inorganic Chemistry (Faculty of Sciences) and Materials Institute, University of Alicante, E-03080 Alicante, Spain; monica_rufete@hotmail.com

* Correspondence: mcroman@ua.es; Tel.: +34-965903975

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Abstract: Monitoring hydrogen consumption has allowed studying the progress of the liquid phase hydrogenation of cyclohexene in toluene with Rh SILP (supported ionic liquid phase) catalysts prepared by the immobilization of the $[\{\text{RhCl}(\text{cod})\}_2]$ complex on different carbon materials. An excess of hydrogen consumption with respect to the required amount for cyclohexene hydrogenation was registered and related with the solvent (toluene) hydrogenation. The study carried out led to unraveling the extent of toluene hydrogenation and to determining if the rate of this reaction is affected by the properties of the carbon material used as support. The results revealed that the Rh SILP catalysts we prepared showed acceptable toluene conversion, with 100% selectivity to the total hydrogenated product, and that the effect of the carbon support is the same as for cyclohexene hydrogenation.

Keywords: $[\{\text{RhCl}(\text{cod})\}_2]$ complex; SILP catalysts; hydrogen consumption; toluene hydrogenation; cyclohexene hydrogenation

1. Introduction

Toluene is a solvent commonly used in several catalytic processes. Examples of its use in the hydroformylation and hydrogenation of olefins can be found in the following references [1–7]. The purpose of most of these research works is to study the transformation of the target substrate into the desired products, usually in relation with the catalysts' properties and the reaction variables. The potentially occurring minor process of solvent hydrogenation remains often unobserved, and it is not taken into account, although it could affect the processes efficiencies. Thus, most studies dealing with hydrogenation reactions in toluene, among them those indicated above [1–7], do not mention the possibility of toluene hydrogenation. This is, likely, because the relatively mild reaction conditions used allow considering negligible the hydrogenation of an aromatic compound, and particularly, toluene as it is much more difficult to hydrogenate than other aromatics, such as naphthalene [8].

Cyclohexene hydrogenation in toluene using SILP (supported ionic liquid phase) catalysts based on carbon materials and the $[\{\text{RhCl}(\text{cod})\}_2]$ complex (chloro(1,5-cyclooctadiene) rhodium(I) dimer, abbreviated as RhCOD) was reported in a previous work [9], which was focused on studying the effect of the carbon support properties on the catalysts' performances. The SILP methodology consists in the immobilization of homogeneous catalysts on solid supports by mean of a solution of the active species in an ionic liquid (IL) that is deposited on their surface [10–14]. This system is a solid heterogeneous catalyst, but the active sites are molecular species able to act as homogeneous catalysts, and thus, it combines the advantages of both, homogeneous and heterogeneous catalysis. Because of that it can be regarded as a hybrid catalyst [10–12,15].

Ionic liquids are being broadly used in the field of catalysis as they can either, act as solvents, constitute SILP systems, or take part of other systems formed by metal nanoparticles and ionic liquids. A recent review shows the achievements of this last type of catalyst in arene hydrogenation [16].

In the aforementioned previous work [9], activity data on the hydrogenation of cyclohexene in toluene (5% vol cyclohexene (CH) in toluene, 333 K, 10 bar H₂, 5 h) were mainly obtained by GC analysis of the solution resulting after reaction. In some experiments, the reaction progress was monitored by the continuous measurement of hydrogen consumption, and a certain excess of consumed hydrogen with respect to the amount required for the actual cyclohexene conversion was detected. Such an observation led us to consider that toluene hydrogenation could also be taking place, and because of that, it was considered interesting to perform a thorough analysis of the whole set of hydrogen consumption versus time curves in order to study, in more detail, this side process.

Besides, toluene hydrogenation is a relevant reaction to both the petrochemical and fine chemical industries [17], and it is a representative compound to study the hydrogenation of aromatics. Important industrial catalysts for the hydrogenation of aromatics, and in particular for toluene hydrogenation, are sulfided Ni-W/Al₂O₃ catalysts [17]. Thus, many research works are focused on Ni catalysts, using more acidic oxide supports (SiO₂-Al₂O₃, TiO₂, zeolites) or some modification with other elements [17–20]. However, in all these cases the catalysts have a high metal loading, usually above 20%, and the reaction conditions involve temperatures and pressures usually above 200 °C and 20 bar, respectively. With the purpose of exploring alternative, and maybe more convenient catalysts, it was considered interesting to determine the capabilities of the [RhCl(cod)]₂ SILP catalysts investigated in this study for such a reaction.

2. Results and Discussion

2.1. Summary of the Supports Characterization Data

The N₂ adsorption isotherms of carbons KA and GeA are type I according to the IUPAC classification, typical of microporous materials, while those obtained for samples SA and T are type I+IV, characteristic of solids with micro and mesoporosity [21] (Figure S1, Supplementary Material). The TPD (temperature programmed desorption) profiles of these samples (Figure S2, Supplementary Material) reveal important differences between them regarding their surface chemistry. The determined textural and chemical parameters were previously reported [9] and are collected in Table S1 (Supplementary Material), where it can be seen that the carbon materials used as supports have a quite different porosity and surface chemistry, both in terms of amount and type of surface oxygen groups.

2.2. Cyclohexene Hydrogenation in Toluene with Homogeneous and Biphasic Systems

Figure 1 shows the hydrogen consumption versus time curves recorded during the catalytic tests carried out with the homogeneous catalyst and the biphasic system. It can be observed that the hydrogen consumption steadily increases in both cases. Considering the molar stoichiometry for cyclohexene hydrogenation (1:1, H₂:CH), 120.6 cm³ is the volume of hydrogen required for a 100% substrate conversion. The data of Figure 1 indicate that hydrogen is consumed in excess with respect to the amount mentioned, and thus, it can be considered that toluene is also being hydrogenated.

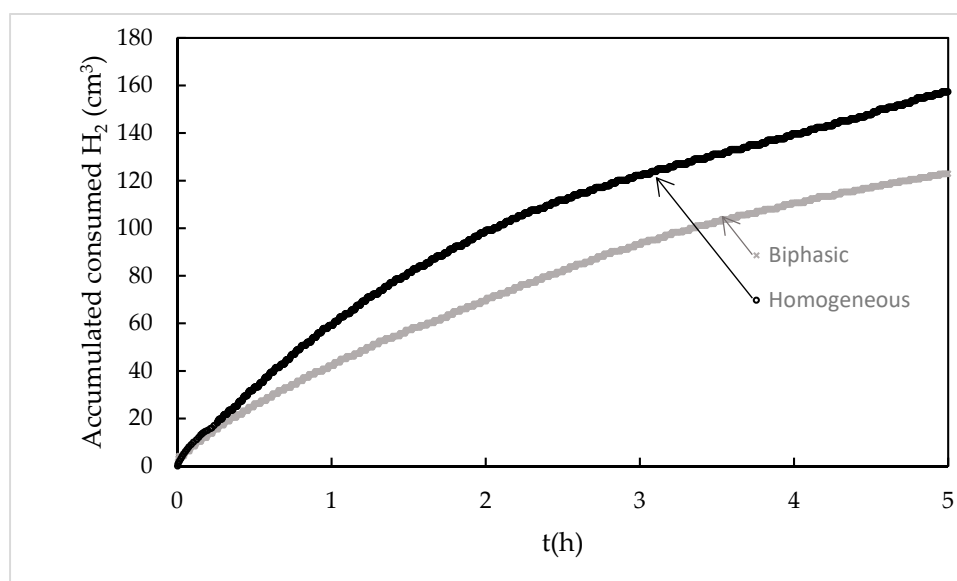


Figure 1. Hydrogen consumption versus time during CH hydrogenation in toluene (5% volume CH in toluene, 333 K, 10 bar H₂) with homogenous and biphasic systems.

Table 1 collects data on the amount of hydrogen consumed at 5 h; cyclohexene conversion at the same time, determined by gas chromatography; and the calculated amount of hydrogen required to achieve such a cyclohexene conversion.

Table 1. Hydrogen consumption, CH conversion and H₂ consumption calculated for the CH conversion measured (5% vol CH in toluene, 333 K, 10 bar H₂, 5 h).

Catalyst	H ₂ cons (cm ³) ^a	CH conv (%) ^b	H ₂ for CH conv (cm ³) ^c	r ₀ (m ³ h ⁻¹) ^d
Homogeneous	142.4	98.4	118.7	27.1
Biphasic	122.8	98.1	118.3	22.8

^a At t = 5 h. ^b Cyclohexene (CH) conversion determined by gas chromatography. ^c H₂ amount calculated to be required for the CH conversion determined by GC. ^d Initial reaction rate calculated as the initial slope (up to 0.5 h) of the curves shown in Figure 1.

Cyclohexene conversion at 5 h is above 98% for both the homogeneous and the biphasic systems. The amount of hydrogen consumed in excess with respect to the amount needed for such a conversion (23.7 and 4.5 cm³ in the homogeneous and biphasic systems, respectively) should be attributed to toluene hydrogenation. Thus, according to these results, toluene conversion in the homogeneous system is higher than in the biphasic one (0.36% versus 0.07%), which can be explained by the more effective contact between the substrate and the Rh complex in the first case.

The initial reaction rate has been estimated by means of the initial slope (up to 0.5 h) of the curves shown in Figure 1 and the calculated data are included in Table 2.

2.3. Cyclohexene Hydrogenation in Toluene with RhCOD SILP Catalysts

Table 2 (column 1) shows the amount of hydrogen consumed at 5 h determined from the H₂ consumption versus time curves obtained in the catalytic tests performed with the Rh catalysts. Figure 2 shows, as an example, the hydrogen consumption versus time curve obtained with catalyst SA20-RhCOD.

Table 2. H₂ consumption and conversion data (5% vol CH in toluene, 333 K, 10 bar H₂, 5 h).

Catalyst	1 H ₂ cons (cm ³) ^a	2 CH _{conv} (GC) (%) ^b	3 H ₂ for CH _{conv} (cm ³) ^c	4 Tol _{conv} from H ₂ excess (%) ^d	5 Tol _{conv} (GC) (%) ^e	6 mol Tol /mol CH	7 r ₀ (m ³ h ⁻¹) ^f
KA20-RhCOD	117.9	52.9	61.3	0.9	0.9	0.3	24.4
GeA20-RhCOD	193.7	92.8	111.9	1.	1.4	0.3	42.6
SA20-RhCOD	225.9	100.0	120.6	1.6	1.7	0.3	109.2
T20-RhCOD	212.8	100.0	120.6	1.4	1.5	0.3	38.6
KA-RhCOD	63.5	16.3	19.6	0.7	0.8	0.9	13.0

^a At t = 5 h. ^b Cyclohexane conversion determined by gas chromatography [9]. ^c H₂ required for the CH conversion. ^d H₂ excess: total hydrogen consumed minus hydrogen required for CH hydrogenation. ^e Toluene conversion determined by gas chromatography. ^f Initial reaction rate calculated as the initial slope (up to 0.5 h) of the H₂ consumption curves.

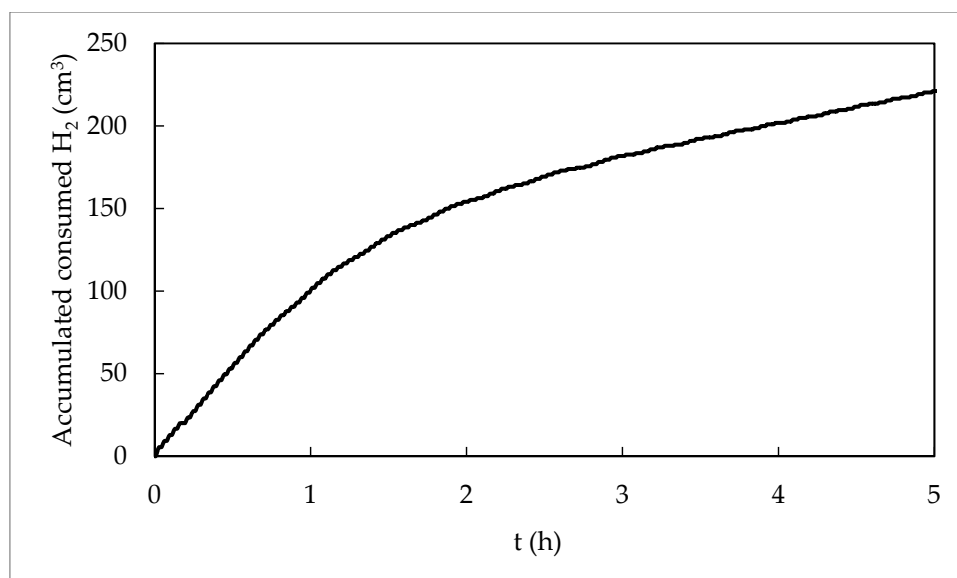


Figure 2. H₂ consumption versus time during CH hydrogenation in toluene with catalyst SA20-RhCOD (5% vol CH in toluene, 333 K, 10 bar H₂).

Cyclohexene conversion at 5 h determined by gas chromatography (data previously reported [9]) is also presented in Table 2 (column 2) along with the amount of hydrogen consumed for such a conversion assuming the 1:1 molar stoichiometry of the CH hydrogenation reaction (column 3). These data show that in all cases, the amount of hydrogen consumed is clearly higher than that required for the cyclohexene conversion measured, and as commented above, the hydrogenation of toluene should be assumed to be occurring.

In order to study such a reaction independently, catalyst SA20-RhCOD was used in a 10 h catalytic test in the same conditions but without adding cyclohexene to the reaction media. The hydrogen consumption we recorded versus time is depicted in Figure 3, where it can be observed that the hydrogen consumption proceeds quite linearly with time.

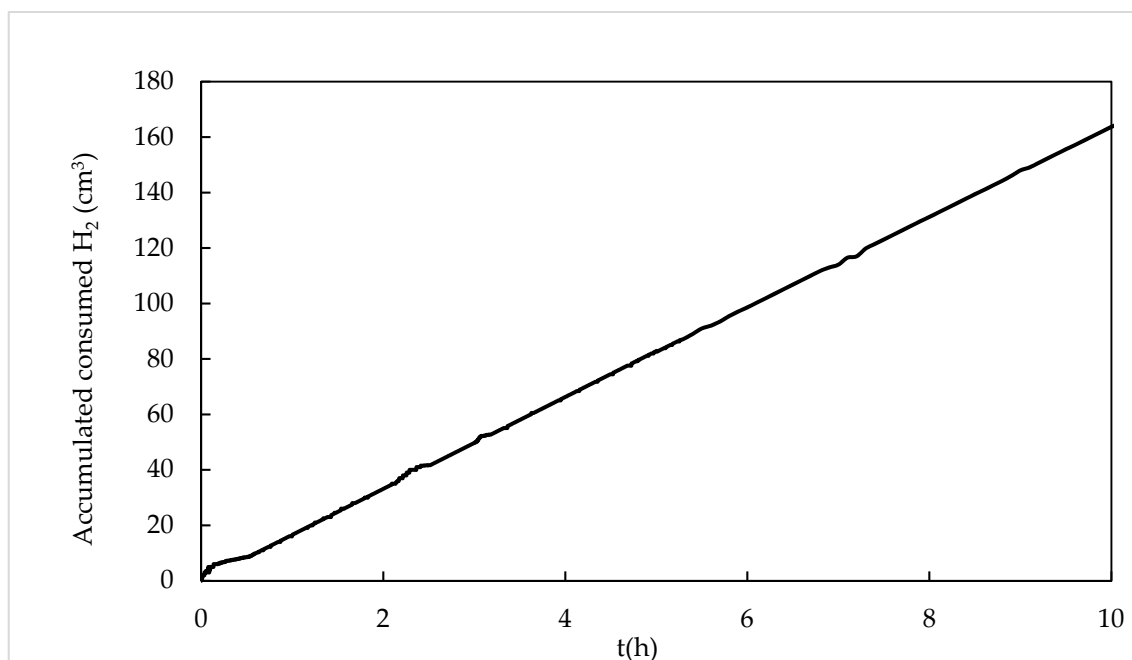


Figure 3. H₂ consumption versus time in the hydrogenation of toluene with catalyst SA20-RhCOD (333 K, 10 bar H₂).

The reaction solution was analyzed by gas chromatography, and methylcyclohexane was the only product observed. Quantification of this product allowed determining that after 10 h reaction time, toluene conversion was 2.30%, which corresponds to 0.0020 mol converted toluene. On the other hand, considering that the amount of hydrogen consumed in 10 h (165.57 cm³, see Figure 3) is used to hydrogenate toluene to methylcyclohexane, the toluene conversion would be 2.04%. Comparison of these two conversion values indicates that there is a good agreement between the results obtained by H₂ consumption and by GC analysis of the solution, supporting the reliability of the measurements done with the device developed by the research group to monitor the reaction by means of the H₂ consumption (see section 3.3.).

As reported in the previous work [9] (and shown in the Supplementary Material), Rh leaching is very low in all SILP catalysts, less than 0.7% (Table S2), and XPS results (Table S3) have shown that Rh in the used catalysts remains in the original electronic state (that is, no reduction takes place upon use), even after three consecutive reaction runs.

The activity of the SA20-RhCOD catalyst for toluene hydrogenation is in line with that reported for SCILL catalysts (solid catalysts with ionic liquid layer) in which dendrimer encapsulated Ru nanoparticles supported on silica are coated with an ionic liquid layer [22]. The catalysts, tested in the liquid phase hydrogenation of toluene at 383 K and 30 bar H₂, show reaction rate values that range from $0.5 \cdot 10^8$ to $2.0 \cdot 10^8$ (mol_{tol}) · (mol_{Ru})⁻¹ m⁻³ h⁻¹ depending on the IL used. The toluene conversion of catalyst SA20-RhCOD in 10 h corresponds to a reaction rate of $1.37 \cdot 10^8$ (mol_{tol}) · (mol_{Rh})⁻¹ m⁻³ h⁻¹ (and to a turn over frequency (TOF) of 0.0190 s⁻¹) which was achieved in significantly milder conditions, 333 K and 10 bar H₂ than those reported in [22]. This catalyst is also clearly more active for toluene hydrogenation than other catalysts based on Rh nanoparticles in ionic liquid [23]. In this case, the authors report TON (turnover number) values (calculated as mol of consumed H₂ per mol of introduced Rh) of 255 and 300 (under reaction conditions of 353 K, 40 bar, 15 h), while the analogous value obtained for the SA20-RhCOD catalyst was 2550. Additionally, in the work of Suppino et al. [24], in which Pt, Pd and Ru catalysts supported on alumina have been developed for the hydrogenation of toluene at 373 K and 50 bar, the reported reaction rates are lower than those reported in the present work. The maximum reaction rate values reported by these authors are about $3 \cdot 10^7$ (mol_{tol}) · (mol_M)⁻¹ m⁻³ h⁻¹.

Once assured that the SILP catalysts are able to hydrogenate toluene at 333 K and 10 bar H₂, the hydrogen consumed in excess during the cyclohexene hydrogenation tests has been attributed

definitively to toluene hydrogenation. This allows calculating toluene (Tol) conversion in each test. Table 2 (column 4) shows toluene conversion calculated in this way. It can be seen that toluene conversion in 5 h with the SILP catalysts ranges from 0.9% to 1.6% and it is bit lower with the KA-RhCOD sample (the one prepared without ionic liquid). Besides, a careful analysis of the GC data led to the determination of methylcyclohexane and to calculate also toluene conversion from these data (Table 2, column 5). Both values (columns 4 and 5) are coherent, and thus, it can be concluded that the hydrogen consumption curves are very useful in order to determine the extent of toluene hydrogenation when it is used as a solvent.

To compare cyclohexene (CH) and toluene (Tol) hydrogenation activities of each catalyst, the molar ratio of converted CH to converted Tol has been calculated and included in Table 2 (column 6). It can be observed that the four Rh SILP catalysts are significantly more active for CH than for Tol hydrogenation (activity for CH hydrogenation is about three times the activity for Tol hydrogenation). However, catalyst KA-RhCOD, prepared without ionic liquid, that is less active than the SILP catalysts in both hydrogenation reactions, is, in relative terms, more active for toluene hydrogenation, bringing the activity for both reactions closer. This result reveals the importance of the ionic liquid to keeping a suitable state of the active species. On the other hand, and as previously reported [9], the Rh SILP catalysts prepared with supports that have micro and mesopores are more active for CH hydrogenation than the one prepared with the essentially microporous carbon support (carbon KA). The equal molar ratio of converted Tol to converted CH (column 6, Table 2) obtained in this study for all the SILP catalysts means that this applies also for the hydrogenation of toluene. That is, the textural properties of the carbon material used as supports have a similar effect in the activity of the Rh SILP catalysts for the hydrogenation of cyclohexene and the hydrogenation of toluene.

The initial reaction rate, estimated by the initial slope (up to 0.5 h) of the H₂ consumption curves, is also included in Table 2 (column 7).

Table 3 shows the calculated TOF and TON data corresponding to hydrogenation of cyclohexene and toluene in the hydrogenation tests of cyclohexene in toluene.

Table 3. TOF (turn over frequency) and TON (turnover number) values in the hydrogenation of cyclohexene and toluene (5% vol CH in toluene, 333 K, 10 bar H₂), determined after 5 h of reacting.

Catalyst	CH Hydrogenation		Tol Hydrogenation	
	TOF (s ⁻¹)	TON	TOF (s ⁻¹)	TON
KA20-RhCOD	0.050	900	0.016	288
GeA20-RhCOD	0.087	1566	0.023	414
SA20-RhCOD	0.094	1692	0.029	522
T20-RhCOD	0.094	1692	0.025	450
KA-RhCOD	0.015	270	0.013	234

These results can be compared with those reported by Pawelec et al. [8] that studied the simultaneous hydrogenation of toluene and naphthalene in the presence of dibenzothiophene using Pt and Pd catalysts supported on zeolites. The process was studied in a continuous down flow catalytic reactor at 448 K and 50 bar. Most of the catalysts studied by these authors show TOF values (calculated from the reported reaction rates) that range from 0.01 and 0.06 s⁻¹. In spite of the differences in the process performance, it can be concluded that the activity of the catalysts studied in the present work for toluene hydrogenation (in the presence of other substrates) is reasonably high, mainly considering the mild operation conditions.

2.4. Kinetics of the Hydrogenation of Cyclohexene in Toluene

Assuming that toluene hydrogenation occurs in a similar way when it is pure and mixed with cyclohexene, the hydrogen consumption registered in the test carried out with pure toluene using sample SA20-RhCOD (Figure 3) was subtracted from analogous data determined with the 5 vol% solution of CH in Tol (Figure 2). The resulting plot, shown in Figure 4, should represent the kinetic curve of the independent CH hydrogenation reaction. This is a typical kinetic curve similar to the one

found previously for cyclohexene hydrogenation with homogeneous Rh and Pd amine complexes [25]. Thus, this analysis could indicate that the catalytic behavior of the RhCOD complex heterogenized by the SILP methodology resembles that of the one being used in homogeneous phase, which supports the hybrid character (homogeneous/heterogeneous) of these catalysts.

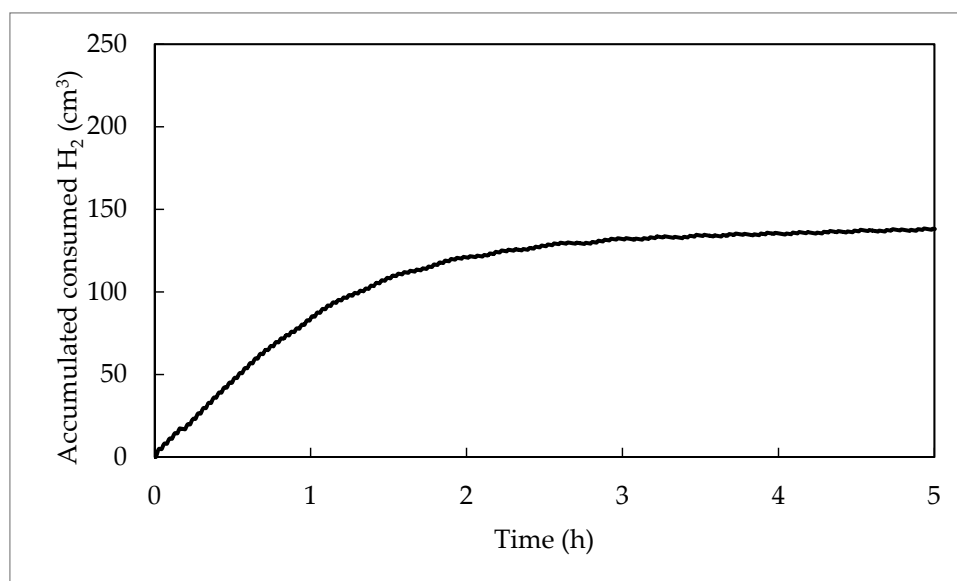


Figure 4. H₂ consumption versus time curve for catalyst SA20-RhCOD obtained by subtracting the H₂ consumption versus time curve determined for pure toluene hydrogenation (Figure 3) from the H₂ consumption versus time curve measured for the hydrogenation of cyclohexene in toluene (333 K, 10 bar H₂) shown in Figure 2.

Additionally, the H₂ consumption versus time curve of Figure 4 shows that the reaction time for a high CH conversion can be reduced. It can be seen that the maximum CH conversion was reached in about 3 h. Thus, the TOF value reported in Table 3 for this sample can be recalculated, considering a 3 instead 5 h reaction time, what brings about a much higher value, 0.157 s⁻¹. Furthermore, if the TOF is calculated for an 80% CH conversion, which is achieved at 1.5 h according to Figure 4, the noticeably high value of 0.250 s⁻¹ is determined.

In summary, monitoring the hydrogenation of CH in toluene by the H₂ consumption occurring during the process has allowed: to determine that toluene hydrogenation also takes place, to show the independent hydrogenation of CH, and to state that the Rh SILP catalysts reported in this work are reasonably active for toluene hydrogenation. Thus, they could be an interesting option for carrying out the hydrogenation of such an aromatic compound.

3. Materials and Methods

3.1. Supports

The SILP catalysts were prepared using as support, the following commercial carbon materials [9]:

GeA—activated carbon of spherical morphology prepared by Gun-ei Chemical Industry (Takasaki-shi, Gunma, Japan,) from a phenolic resin. The average diameter of the spheres is 150 μm.

KA—activated carbon of spherical morphology produced by Kureha Corporation (Takasaki-shi, Gunma, Japan) from petroleum pitch. The average diameter of the spheres is 780 μm.

T—carbon black T-10157 produced by Columbian Chemical Company (Centerville, LA, USA).

SA—powder activated carbon named SA-30 produced by MeadWestvaco (Richmond, VA, USA).

They were characterized by gas adsorption (N_2 at 77 K and CO_2 at 273 K) to determine the textural properties, and by temperature programmed desorption (TPD) experiments to assess the surface chemistry, as indicated in [9].

3.2. Catalysts' Preparations and Characterizations

The SILP catalysts were prepared using the ionic liquid (IL) 1-butyl-3-methyl imidazolium hexafluorophosphate ([bmim][PF₆]) and the $[RhCl(cod)]_2$ complex (chloro(1,5-cyclooctadiene) rhodium(I) dimer (abbreviated as RhCOD)). Preparation of catalysts with 20 wt% IL and 1 wt% Rh loading was carried out inside a Schlenk apparatus as follows: the appropriate amounts of ionic liquid and RhCOD and 2 cm³ anhydrous acetone, were added under Ar atmosphere to 0.5 g of the support (previously outgassed under vacuum at 423 K for 5 h). The mixture was stirred until the sample had a dry appearance. The catalysts generic name is SUP20-RhCOD, where SUP is in each case the support name (see above, indicated in Section 3.1) and 20 refers to the weight percentage of ionic liquid in the sample. A catalyst without ionic liquid was prepared using carbon KA as support by the same procedure but skipping the addition of the [bmim][PF₆] compound. This catalyst was named KA-RhCOD.

The amount of Rh in fresh and used catalysts was determined by ICP, after decomposition of the carbon material by acid treatment as indicated in [9]. Determinations on the fresh catalysts revealed Rh loading between 0.95 and 0.98 wt%. Fresh and used catalysts were analyzed by X-ray photoelectron spectroscopy (XPS) using a VG-Microtech Multilab 3000 apparatus (Thermo Fisher Scientific, Waltham, MA, USA) in order to determine the chemical state of rhodium, and by transmission electron microscopy (TEM) using a JEM-2010 JEOL microscope (JEOL, Akishima, Japan) to obtain information about the samples' morphologies and compositions.

3.3. Liquid Phase Hydrogenation of Cyclohexene in Toluene

In a typical test, 30 mg catalyst (2.9 μ mol Rh) and 10 cm³ of a 5 vol% cyclohexene in toluene solution ($4.9 \cdot 10^{-3}$ mol cyclohexene) were introduced in the stainless steel reactor. After closure, a leak test was performed, and then, the system was heated to 333 K and pressurized with 10 bar H₂. In a reference biphasic system, the catalyst consisted of a solution of 1.5 mg RhCOD in 0.01 cm³ ionic liquid. For a homogeneous phase test, 1.5 mg of RhCOD was dissolved in the substrate solution; 1.5 mg is the amount of Rh(COD) contained in 30 mg of the hybrid catalysts. The reaction time was 5 h and its evolution was followed through the hydrogen consumption by means of an experimental device specifically developed for this purpose. It is based on the continuous measurement of the pressure inside the reactor and the automatic activation of a hydrogen feeding system in case the pressure falls below a certain value (10 bar in this case). Thus, when the pressure decreases (below 10 bar) due to the reaction progress, a valve automatically opens to feed hydrogen into the reactor until the initial fixed pressure is reached again. A mass flow meter allows registering the amount of hydrogen introduced into the system during the reaction and the data accumulated is the hydrogen consumed. A scheme of the experimental device is shown in Figure 5.

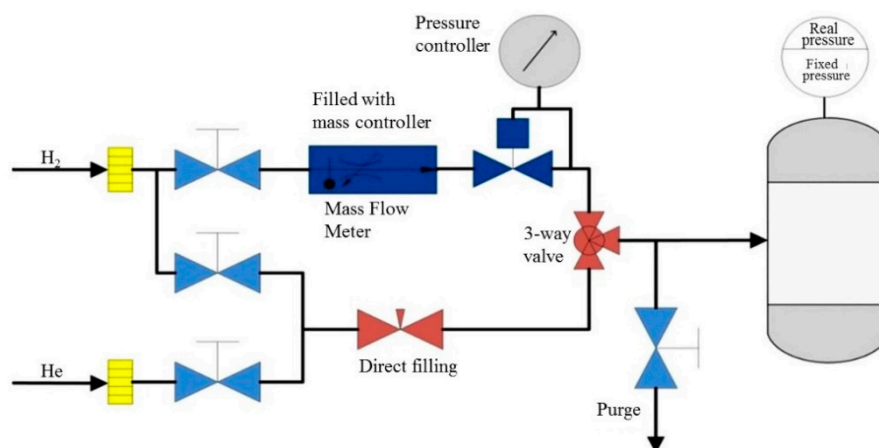


Figure 5. Scheme of the hydrogen consumption monitoring system.

The absence of leaks was tested prior each experiment by filling the system with a certain gas pressure and ensuring that there was not a pressure drop.

On the other hand, the reaction medium was analyzed by gas chromatography (HP6890 Series II with the column HP-1 MethylSiloxane 30 m × 250 μm × 0.25 μm, Palo Alto, CA, USA). For the analysis, 150 μL of the problem solution was mixed with 250 μL acetone and 100 μL of a 7285 ppm solution of decane (internal standard) in acetone.

4. Conclusions

Following the progress of the hydrogenation of cyclohexene in toluene by hydrogen consumption has allowed unraveling the fact that SILP catalysts are able to hydrogenate toluene in liquid phase (at 333 K and 10 bar H₂) with 100% selectivity to methylcyclohexane. Analysis of the conversion versus time curves has permitted us to determine the extent of toluene conversion with the studied catalysts. Additionally, such an analysis is useful to unveil the kinetic curve of the independent hydrogenation of cyclohexene in the cyclohexene/toluene reaction media.

Rh SILP catalysts prepared with carbon materials that have both, micro and mesopores are more active in the hydrogenation of toluene than those prepared with essentially microporous carbon materials. These results are in agreement with the previously reported behavior of these catalysts for cyclohexene hydrogenation.

It must be also pointed out that the prepared carbon based SILP catalysts are active for the hydrogenation of toluene, showing activity values which are comparable with those found in the literature.

Supplementary Materials: The following are available online at www.mdpi.com/xxx/s1. Figure S1: N₂ adsorption isotherms at 77 K of the carbon materials. Figure S2: TPD profiles (CO₂ (empty markers) and CO (filled markers refer to the secondary axis)) of the original carbon materials (experimental conditions: 10 Kmin⁻¹ up to 1273 K and 100 cm³min⁻¹ He). Table S1: Main characterization of the carbon materials used as supports: Textural properties and oxygen surface chemistry. Table S2: Rh leaching upon catalysts use (5 vol% cyclohexene in toluene, 333 K, 10 bar H₂, 5 h). Table S2: XPS Rh 3d binding energy data of fresh and used catalysts (5 vol% cyclohexene in toluene, 333 K, 10 bar H₂, 1.5 h).

Author Contributions: conceptualization, M.C.R.-M.; methodology, M.R.-B. and M.C.R.-M.; formal analysis, M.R.-B. and M.C.R.-M.; investigation, M.R.-B. and M.C.R.-M.; data curation, M.R.-B. and M.C.R.-M.; writing—original draft preparation, M.R.-B. and M.C.R.-M.; writing—review and editing, M.C.R.-M.; supervision, M.C.R.-M.; project administration, M.C.R.-M.; funding acquisition, M.C.R.-M.

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Conflicts of Interest: The authors declare no conflict of interest.

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