

Influence of the CO Adsorption Environments on its Reactivity with (111) Terrace Sites in Stepped Pt Electrodes under Alkaline Media

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Abstract

The effect of the electrode potential in the reactivity of platinum stepped single crystal electrodes with (111) terraces towards CO oxidation has been studied. It is found that the CO adlayer is significantly affected by the potential at which the adlayer is formed. The electrochemical and FTIR experiments show that adsorbed CO layer formed in acidic solution at ~ 0.03 V vs. SHE is different from that formed at -0.67 V vs. SHE in alkaline solutions. The major effect of the electrode potential is a change in the long-range structure of CO adlayer. The adlayer formed in alkaline media presents a higher number of defects. These differences affect the onset and peak potential for CO stripping experiments. The stripping voltammogram for the adlayer formed at -0.67 V vs. SHE always shows a pre-wave and the peak potential is more negative than that observed for the adlayer formed at 0.03 V vs. SHE. This means that the apparent higher activity for CO oxidation observed in alkaline media is a consequence of the different CO adlayer structure on the (111) terrace, and not a true catalytic effect. The different behavior is discussed in terms of the different mobility of CO observed depending on the electrode potential. Also, the FTIR frequencies are used to estimate the pzc (potential of zero charge) for the Pt(111) electrode covered with a CO adlayer.

Keywords: Electrocatalysis; carbon monoxide; stepped Pt electrode; alkaline media, pzc.

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1. Introduction

In the field of electrocatalysis, the reactivity of platinum-based electrodes for the carbon monoxide oxidation reaction strongly depends on the history of the electrode surface as well as the applied potential at which the CO adlayer was formed. This phenomenon was firstly reported by Kita *et al.* in 1988.¹ They observed that the voltammetric profile for CO stripping on a polycrystalline platinum electrode always exhibited a pre-oxidation wave that preceded the main CO electro-oxidation peak when electrode potentials for CO adsorption was held under the hydrogen adsorption/desorption region. On the other hand, when the electrode potentials for CO adsorption remained within the so-called double layer region, the stripping voltammetry corresponding to this CO adlayer did not exhibit the pre-oxidation wave and occurred in a single peak. Recently, ATR-SIERS experiments indicate these differences are related to changes in the water-CO structure obtained at different dosing potentials.²

Although the physical origin (at a molecular level) of such changes are not known yet, it was soon realized that the potential for CO adsorption on platinum polycrystalline electrodes had a close relationship with the maximum CO coverage (θ_{CO}) obtained on the electrode surfaces and with the kinetics of CO oxidation.³⁻⁵ In this way, the experimental conditions to reach the maximum CO surface coverage require the adsorption at a potential within the hydrogen underpotential deposition region (H_{upd}), i.e., in potential range where hydrogen is adsorbed on the surface but there is no hydrogen evolution. This type of behavior is better studied and understood with platinum single crystal electrodes, since the surface structure is well known and in, some cases, ordered CO structures are formed upon adsorption. This is the case for the well-oriented Pt(111) electrode, for which two different structures were observed by *in situ* STM in the presence of CO in solution depending on the potential for CO adsorption;⁶ namely a $p(2\times 2)-3CO$ with $\theta_{CO} = 0.75$ formed at ~ 0.00 V *vs.* RHE, or a $(\sqrt{19}\times\sqrt{19})R23.4^\circ-13CO$ with $\theta_{CO} = 0.69$ at 0.25 V. It is clear that none of

those structures can be formed at a polycrystalline surface because long-range (111) domains are required to form those compressed and ordered CO structures.⁷ However, the knowledge gained with the single crystal electrodes is useful to understand the origin of preoxidation wave at a molecular level. The potential at which the CO adlayer was formed has a clear relationship with the CO coverage reached at those potentials and consequently with the CO oxidation kinetics.⁸⁻⁹ Regarding the oxidation mechanism, it was clear from the beginning that the CO oxidation reaction required the presence of adsorbed oxygenated species.¹⁰ The studies carried out with single crystal electrodes in acidic solution seem to indicate that the process occurs through the Langmuir-Hinshelwood mean field mechanism,¹¹⁻¹² which suggested that CO diffusion on the surface was fast.¹³⁻¹⁴

The changes in the oxidation of CO depending on the adsorption potential in the SHE scale, are very clear when the voltammetric profile for the CO oxidation reaction in alkaline solutions is compared to that observed in acidic media. These differences were demonstrated by Koper and his co-workers for stepped surfaces containing (111) terrace sites.¹⁵⁻¹⁹ In alkaline media, several peaks appear for the oxidation of CO on those surfaces and a prewave is always present. In general, two main peaks (aside from the pre-wave) are observed in the voltammetry for stepped surfaces with (111) terraces. The one at higher potential appears in the same potential regions as the main peak for the Pt(111) electrode. The charge for the peak at lower potential is proportional to the step density and its peak potential depends on the site symmetry.¹⁵ These facts allow assigning each peak to a different surface site. That appearing at lower potentials was initially ascribed to the oxidation of CO adsorbed on low coordination sites (such as kink/step sites) because those sites can act as a center for nucleation of the oxygen-containing species that are required for CO oxidation. However, additional experiments indicated that the peak at low potentials corresponded

to the CO oxidation adsorbed on terrace sites close to the step.^{16, 20} The peak at higher potentials is then related to the oxidation of CO on the remaining terrace sites and on the step sites.

From the studies with stepped surfaces vicinal to the (111) pole, e.g., having (111) terraces, it can be considered that two different sites are present in a stepped surface when the CO oxidation reaction is considered: those very close to the step and the rest of the sites. The rate constant for the CO oxidation process in each site is potential dependent (following Butler-Volmer equation) and it is significantly higher for the sites close to the step for the same electrode potential. The CO oxidation process on each type of site ignites when the rate constant reaches a threshold value. At given potential, the rate constant for the sites close to the step will reach the threshold value and the CO molecules on those sites will be immediately oxidized. Then, two different situations appear depending on the pH. At low pH, the experimental evidence suggests that CO diffusion rate is fast (significantly faster than threshold value for the CO oxidation rate).^{9, 11-12} CO molecules adsorbed on sites far from the step, whose CO oxidation rate constant is still below the threshold value for ignition, will immediately diffuse to the sites close to the step, where they are readily oxidized. These conditions give rise to only one peak in the voltammogram for the CO oxidation process, as has been experimentally observed.⁹ This voltammetric peak will have the contribution from the whole CO layer, but the oxidation has occurred exclusively on the sites close to the step. On the other hand, it is proposed that the CO diffusion rate in alkaline media is very slow.²¹ When the CO molecules adsorbed on sites close to the step are oxidized, the flow of CO molecules from sites far from the step reaching the activated sites is very small. To reach a significant oxidation rate for those molecules far from the step sites, the potential has to be increased so that the CO oxidation rate constant on remaining sites reaches the threshold value. Two voltammetric peaks are then obtained, one at low potentials with the contributions from the sites close to the step (and some

from sites far from the step) and the one at higher potentials which contains the charge from oxidation of the remaining molecules.

Another important difference in the behavior of the alkaline solutions with respect to acidic media is that the prewave is usually present. The appearance of this prewave has been associated to the presence of defects in the CO adlayer which make the initial stages of CO oxidation easier. In this respect, Spendelow *et al.*²² showed that when a CO adlayer was formed on a Pt(111) electrode under a 0.1 M NaOH environment at 0.1 V *vs.* RHE and then transferred to a 0.5 M H₂SO₄ solution, the corresponding voltammetric CO stripping always presented a preoxidation peak. In the reverse experiment, when the CO was adsorbed in acidic solution at 0.1 V *vs.* RHE and transferred to alkaline solutions, the prewave was absent. These experiments clearly suggest that the electrode potential has an important effect on the CO structure. In this respect, it has been shown that the $p(2\times 2)-3\text{CO}$ structure, observed when CO is present on the solution, can be obtained both in acidic and alkaline solutions.²³ However, there is a significant change in the size of the ordered domains, which are smaller in alkaline solutions.²³⁻²⁴ Since defects concentrate in the borders of the different ordered domains, the smaller the domains in the CO layer are, the larger amount of defects such layer contains. Incidentally, this interpretation supports the lower mobility of CO on Pt(111) electrodes in alkaline environments as compared to acidic media. Regarding the kinetics of the initial stages of the oxidation processes, it was clear from the beginning that they did not followed Langmuir- Hinshelwood kinetics, which lead to suggest that the oxidation mechanism for those initial stages followed Eley-Rideal mechanism.¹¹ This oxidation mechanism has also been proposed for the oxidation of the prewave in alkaline solutions.²⁵

All the previous results indicate that the electrode potential can have an important effect on the CO adlayer structure and its oxidation kinetics. To explore this effect, a detailed study of CO oxidation on stepped Pt(s) [($n-1$)(111) \times (110)] electrodes has been carried out in alkaline and acidic

media. We will show that the structure of the CO adlayer with (111) terrace sites in stepped Pt electrodes is strongly influenced by the solution pH at which the CO adlayer was formed, whereas the reactivity for CO oxidation in both media is very similar, provided that the CO adlayer have the same characteristics.

2. Experimental

Platinum single crystal electrodes were prepared according to Clavilier technique,²⁶ flame annealed and cooled down in a H₂ + Ar controlled atmosphere, which assures that the surface topography is in agreement with the nominal values.²⁷ The single crystal surfaces used in this work were the Pt(111) electrode and its vicinal surfaces containing (110) monoatomic steps. The notation of these electrodes is Pt(*n*+1, *n*+1, *n*-1) ≡ Pt(s) [(*n*-1)(111)×(110)], where *n*-1 denotes the number of atomic rows on the (111) terrace parallel to the (110) monoatomic steps. The electrode areas are between 3 and 5 mm² for the electrodes used in cyclic voltammetry and chronoamperometry, whereas in IR experiments the areas are ca. 11-13 mm². Experimental details are given elsewhere²⁰, namely, a three-electrode cell equipped with a platinum counter electrode and reversible hydrogen electrode (RHE) was used. All the electrode potentials in this work are quoted vs. RHE unless otherwise stated. It should be noted that the RHE scale is a pH dependent scale and it is useful for studying processes whose equilibrium potential is also pH dependent, such as CO oxidation. Thus, the RHE scale shifts with pH in the same way that the equilibrium potential for the studied processes, which facilitates the interpretation of the results. However, other important parameters in the interfacial properties, i.e., the interfacial charge, are dependent on the standard thermodynamic scale in electrochemistry, the SHE scale. Thus to study the effects of the electrode potential and the interfacial charge in the process, the values in the RHE scale should be converted

in the SHE scale. To avoid misunderstandings between both scales, we will term that potential as the SHE potential.

The electrode potentials were controlled using a waveform generator (EG&G PARC 175) together with a potentiostat (Amel 551) and a digital recorder (eDAC ED 401). Solutions were prepared using NaOH (Merck KGaA, 99.99 %), H₂SO₄ (Merck, Suprapur), HClO₄ (Merck, suprapur) and ultrapure water (18.2 MΩ cm). Ar (N50, Air Liquide) was used for deoxygenating the solutions. CO (N50, Air Liquide) adlayers were always formed at 0.1 V from saturated solutions. After forming the adlayer, CO was purged from the solution by bubbling Ar for 15 min, and the CO stripping voltammogram is recorded in the same solution. In some cases, CO was dosed at 0.1 V in a H₂SO₄ solution, transferred to a cell containing a 0.1 M NaOH solution, immersed at 0.1 V, and finally recording the CO stripping voltammogram in this latter solution. Charge densities were corrected for the adsorption of anions.²⁸

In situ FTIRS spectra were registered using a Nicolet (Model 8700) spectrometer, equipped with a MCT detector. The spectroelectrochemical cell was equipped with a CaF₂ prism beveled at 60°. ²⁹⁻³⁰ One hundred interferograms with a resolution of 8 cm⁻¹ were collected for each spectrum using *p*-polarized light. Spectra was acquired in a 25 mV sampling interval from 0.05 V to 0.70 V. Reference spectra was collected at 0.90 V. At this potential value, the CO adlayer has been completely oxidized. The spectra are presented in absorbance units as $A = -\log[(R_1 - R_2)/R_1]$, where R_2 and R_1 are the reflectance values for the single beam spectra recorded at the sample and the reference potential, respectively. Positive bands in the spectra correspond to species present at the interface at the sample potential, whereas negative bands are linked to species at the interface at the reference potential but not at the sampling potential.

3. Results and discussions.

3.1. Electrochemical characterization of the surfaces in alkaline medium

Cyclic voltammetric profiles for the Pt(111), Pt(776) and Pt(554) electrodes in alkaline medium are displayed in Fig. 1. The characteristic features and the stability of the voltammograms upon cycling imply that the electrolyte solution is free of impurities. For Pt(111) electrode, the small feature at 0.40 V (positive going scan) indicates that electrode surface contains a small amount of defect sites with (100) symmetry as well as a similar amount of (110) defects at 0.26 V.³¹ The integrated charge involved in these features is less than 1% of that related to the hydrogen adsorption on the (111) planes ($\sim 156 \mu\text{C cm}^{-2}$). For the stepped surfaces, the sharp peaks appearing at ~ 0.26 V (in the positive going scan) are associated with adsorption processes on the (110) step sites.³¹ In order to explain the higher activity for CO oxidation of the stepped surfaces in alkaline media, it has been proposed that these peaks correspond to a competitive adsorption process between OH and H, so that the peak in the positive scan direction corresponds to the desorption of H from the step sites and the simultaneous adsorption of OH.³² Moreover, the charge under these peaks is larger than that expected for hydrogen adsorption processes on the step sites, so it is clear that a competitive adsorption processes between hydrogen and additional species give rise to these peaks.³³ Additionally, the complex potential dependence of the peak with the pH suggests that O is coadsorbed with OH in a ratio that depends on the pH.³³

3.2. Effect of adsorption potential and anion adsorption in CO oxidation

Figure 2 shows the cyclic voltammograms for the oxidation of a CO layer adsorbed at 0.1 V on the Pt(111) surface. As can be seen, there is a clear dependence of the peak potential and sharpness on the anion present on the surface and the pH. In general, the onset for CO oxidation follows the trend $\text{NaOH} < \text{HClO}_4 < \text{H}_2\text{SO}_4$. It should be noted that the oxidation peak in 0.1 M

NaOH exhibits a significant prewave, as described previously.¹⁶ Regarding the peak potential differences between perchloric and sulfuric acid solutions, the higher peak potential for sulfuric acid solutions is related to the specific adsorption of sulfate,³⁴ which hinders the adsorption of the required oxygen containing species necessary for the oxidation of CO through the Langmuir-Hinshelwood mechanism.¹¹ On the other hand, OH adsorption can occur easier on the defects present in the CO adlayer. In fact, it has been suggested that the onset potential depends on the number of those defects,³⁵ which may explain the lower onset for the CO oxidation in alkaline solutions. Moreover, the appearance of the prewave has been associated to the presence of defects in the CO adlayer²¹ or CO oxidation on pristine Pt(111) surfaces with initially disordered (111) steps as observed by *in situ* STM.³⁶ This relationship between the prewave and the defects in the adlayer will be considered in the next sections of this manuscript.

When (110) steps are deliberately introduced in the (111) terrace (Figure 3) changes in the onset and peak potential for CO oxidation occurs. In acidic solutions, the CO stripping process takes place in a single peak. As occurs for the Pt(111) electrode, the CO stripping peak potential in 0.5 M H₂SO₄ solution is ca. 0.1 V higher than that measured in 0.1 M HClO₄ solution.³⁴ When compared to the Pt(111) electrode, peak potentials shift to lower values (ca. 0.07 V for the Pt(554) electrode), due to the catalytic effect of the steps in the reaction.^{9, 12}

On the other hand, the CO stripping voltammetric profile splits into multiple peaks in alkaline solutions. Those peaks are associated to the oxidation of CO on the different surface sites (steps and terraces sites). The lowest potential peak is related to the oxidation of CO on (111) terrace sites close to the step, whereas the CO oxidation on the step and on the remaining (111) terrace sites occurs in the peak at the highest potential.^{16, 20} The double assignment of this peak can be corroborated when the stripping peak of a CO step-decorated surface is compared to that obtained for a CO monolayer (figure 3 – green line). CO-decorated steps were obtained by partial CO adlayer

stripping, as described in detail in reference ²⁰. As can be seen, the charge under the peak at ca. 0.79 V for the step-decorated surface is smaller than that measured for fully covered CO surface, indicating that this peak also contains the contribution of the oxidation of CO molecules in terrace sites far from the step line. In fact, the difference in charge between both experiments is affected by the scan rate and the terrace width.^{15, 21} The longer the terrace width and the faster the scan rate are, the larger the difference in charge for this peak between both experiments is.

The peak potential for this latter peak is equal to that measured on the Pt(111) electrode. It is clear, then, that the (110) steps have a positive catalytic effect in the oxidation of CO, since the onset for the oxidation moves towards negative potential values when steps are present on the electrode surface. However, there is a significant difference between acidic and alkaline solutions. As has been already mentioned, the different behavior of the CO oxidation process in acidic or alkaline media is probably related to a different mobility of CO on the surface depending on the interfacial pH.^{15-16, 21} It has been proposed that the diffusion of CO on the surface is very fast in acidic solutions, so that, when CO close to the step is oxidized, these sites are immediately occupied by incoming CO molecules and the whole oxidation process occurs in a single peak. In alkaline solutions, the catalytic effect of the step is only restricted to the CO adsorbed on the (111) terrace sites close to the step. The proposed low diffusion rate for CO in alkaline solutions prevents the arrival of the CO molecules to the newly liberated sites, giving rise to two peaks, one related to the presence of the steps and the second one with the characteristics of the (111) infinite terraces.

These results seem to suggest that alkaline solutions are more active for the oxidation of CO on Pt, since the onset and peak potentials in the RHE scale are lower than those measured in acidic solutions. However, as it will be shown below, this result is mainly a consequence of the differences in the CO adlayer structure on (111) domains obtained in different electrolytic conditions and not a true catalytic effect.

3.3. Effect of the CO adsorption potential on its oxidation

It has been clearly shown that the CO oxidation potential is affected by the potential at which the CO adlayer is formed.³⁻⁵ For that reason, a new experiment has been carried out in which CO is adsorbed at 0.1 V in 0.5 M H₂SO₄ solution and transferred to a cell containing 0.1 M NaOH in which CO is oxidized. In the SHE scale, this means that CO is formed at ca. 0.03 V vs. SHE in the sulfuric acid solution, transferred to the second cell, immersed at ca. -0.67 V vs. SHE and, then, oxidized. For the adlayer formed in NaOH, the adsorption potential is ca. -0.67 V vs. SHE. This way, the comparison between two layers formed at two very different SHE potentials (and interfacial charge) can be carried out. As can be seen in figure 4, the changes in the CO stripping processes are important. For the adlayer formed in NaOH solutions, the adsorbed CO reaction starts at around 0.38 V in the preoxidation wave and finishes at ~0.9 V. In this preoxidation wave, some small peaks are visible, probably related to the oxidation of CO on the small amount of defects present on the surface,³⁶ in a similar way that CO oxidation is activated in the terrace sites close to the steps. These peaks are followed by another oxidation process that has a maximum current density at ~0.79 V. In case of CO adlayer formed in the sulfuric acid solutions, the stripping voltammetry in alkaline solution does not exhibit any pre-oxidation wave and only one peak appears at higher potentials. In both cases, the obtained CO coverages are equal, within the experimental error ($\theta_{\text{CO}} \approx 0.7$), suggesting that the CO pre-oxidation wave cannot be related to a different CO surface coverage and that the changes should be linked to differences in the CO adlayer structure. As aforementioned, the onset for CO oxidation and peak potential strongly depend on the local CO adlayer structure on the terrace domains and on the number of defects of the adlayer. This fact suggests that the adlayer formed at 0.03 V vs. SHE in acidic solutions has a lower amount of structural imperfections without changing the charge associated the adlayer

oxidation, that is, the measured CO coverage. Model calculations have shown that changes in the CO coverage that are well below the accuracy of the charge measurements in the voltammetry (ca. 1%) lead to significant changes in the peak position.³⁵

Similar experiments were carried out with the Pt(776) and Pt(554) electrodes (figures 5 and 6). As aforementioned, the voltammetric profile for CO stripping contains multiple peaks for the Pt(776) electrode (Figure 5). When the previous experiment is compared to the new one, a major change in the CO stripping profile in alkaline media for the layer adsorbed in acidic solution can be observed. As occurred for the Pt(111) electrode, the prewave is absent and the onset of CO oxidation has been displaced towards higher potential values. However, only the potential for first peak shifts significantly towards higher potential values (ca. 120 mV), whereas the second peak only moves ca. 10 mV. It is also evident that the charge density increases for the process at high potentials. This effect can be related to the slower diffusion of the CO_{ads} molecules towards the terrace sites close to the step. As the scan rate for these experiments is the same, the time elapsed between both peaks for the CO adlayer formed in sulfuric acid solutions is shorter, that is, the time from the moment the sites on the terrace close to the step are liberated and the onset for the second peak has been reduced significantly. This way, a lower number of CO molecules away from the step are able to diffuse to sites close to the lower part of the step and being oxidized. A similar effect has been observed when the scan rate for the stripping process is increased.¹⁵ It should be stressed again that the CO coverages for both experiments were equal within the experimental error ($\theta_{\text{CO}} \approx 0.7$).

Finally, the behavior for Pt(554) electrode (Fig. 6) is qualitatively identical to that observed for Pt(776) electrode. The only quantitative differences are related to small changes of the peak potentials and charge, which in turn depend on the step density. The peak related to the CO oxidation on the (111) terraces close to the step site appears only ca. at 20 mV more positive

potentials for the surface with the lower step density, that is the Pt(776). Also, the charge measured in the second peak for the Pt(554) is smaller because the number of sites far from the step is lower.

The comparison between the measured curves measured in NaOH when CO was dosed at 0.1 V in acidic solution and that obtained in 0.1 M HClO₄ (figures 4B and 6B), shows that differences in the onset potential for CO oxidation in both environments are minimal in the RHE scale, that is, the initial stages of adsorbed CO oxidation are similar in both media. It should be highlighted that, in both cases, the CO adlayers were formed at the same SHE potential. Once the CO oxidation process have started, the differences in the oxidation profile are then related to the slow diffusion of CO towards the most active sites in Pt/alkaline solution interface as compared to the Pt/acidic solution interface.

To further analyze the difference in the kinetics of CO oxidation when the CO adlayer was formed at different SHE potentials, chronoamperometric curves for the Pt(554) were recorded at 0.58 V in 0.1 M NaOH (figure 7). The transients measured for CO stripping at 0.58 V in 0.1 M NaOH for the two different adlayers have different time characteristics. For the adlayer formed in 0.1 M NaOH rapid oxidation takes place, giving rise to a peak at ca. 1 s, whereas the adlayer formed in acid solution is more stable and has a peak at much longer times, ca. 26 s. However, as can be seen in insert I₂ of figure 7, the only signal that is blocked after the stripping at that potential value is that associated to the adsorption of hydrogen on the (110) step sites in both cases. A similar behavior was observed for Pt(776) (not shown). This fact implies that CO adsorbed on the terrace sites has been completely oxidized in the transient and the only CO remaining is that adsorbed on the step sites. In a previous work, these transients have been modeled with a nucleation and growth mechanism in which anisotropic effects are included.²¹ This model gives a typical tailing in the current after in the final part of the transient that can be observed in insert I₁ of figure 7 after 1 s for the adlayer formed at 0.1 V in NaOH. For the adlayer formed in sulfuric acid, however, the

onset and peak times are significantly longer, which is clearly associated with a much lower nucleation rate.

3.5. In situ Fourier Transform Infrared Spectroscopy data

These results indicate that the differences in the behavior between acid and alkaline media are only related to the different CO adlayer properties obtained depending on the SHE dosing potential and the different apparent diffusion rates which is also related to the different SHE potentials of the experiments.²¹ Additional characterization of the adlayer structures can be obtained by analyzing the FTIR spectra. From *in situ* FTIR measurements, although it is not possible a quantitative assignment of CO bonding modes at the surface due to the strong energy transfer between CO bands,^{6,37} CO-CO interactions can be easily shown, which in turn are sensitive to the local CO environment.

Figure 8 shows the *in situ* FTIR spectra for adsorbed CO at 0.1 V in different conditions for the Pt(111) electrode, and two stepped electrodes: Pt(17 17 15) \equiv Pt(s)-[16(111) \times (110)] and Pt(443) \equiv Pt(s)-[7(111) \times (110)]. The CO adlayer was formed at 0.1 V in a NaOH solution and analyzed in same electrolyte, adsorbed at 0.1 V in 0.5 M H₂SO₄ environment and analyzed in NaOH solution and adsorbed at 0.1 V in 0.1 M H₂SO₄ solution and analyzed in same acidic solution. (0.1 M H₂SO₄ solution was used instead of 0.5 M H₂SO₄ solution to prevent corrosion of the window prism. However, no significant differences with were observed in experiments between both acidic media.) The reference spectra is always recorded at 0.9 V in the respective solution after complete CO oxidation. For the top spectrum on the Pt(111) electrode, two broad bands at \sim 2025 cm⁻¹ and \sim 1793 cm⁻¹ assigned to the linear (CO^l) and bridge bonded (CO^b) CO, respectively, are clearly visible. These frequencies are in agreement with previous results from Garcia and Koper for Pt(111)/CO in alkaline media.¹⁹ These bands are strongly affected by the dipole-dipole

coupling,³⁷⁻³⁸ and the absorption intensity is transferred towards the higher frequency band at the expense of the lower frequency band. For linear CO band, its corresponding Stark shifts is $dv_{\text{CO}^l}/dE = 45 \text{ cm}^{-1} \text{ V}^{-1}$. The values for the bridge bonded CO are not given due to low intensity of this band, which precludes an accurate determination of its peak position. The Stark shift obtained in this case is higher than that reported previously in the literature for the same system in alkaline media ($37 \text{ cm}^{-1} \text{ V}^{-1}$).¹⁹ Two additional bands appear in the spectra, at lower frequencies. The broad band at $\sim 1630 \text{ cm}^{-1}$ is related to the O-H bending mode of water in the thin layer³⁹ and other band at $\sim 1400 \text{ cm}^{-1}$ is attributed to the carbonate anion, formed at the reference potential after the oxidation of CO.⁴⁰

When the CO adlayer was formed on Pt(111) in H_2SO_4 solution and the electrode covered by CO is transferred to the NaOH solution, where the spectra are acquired, CO^l and CO^b bands appear at $\sim 2029 \text{ cm}^{-1}$ and $\sim 1794 \text{ cm}^{-1}$, respectively, values that are higher than those observed in the previous case. Additionally, these bands are narrower. The higher frequencies and narrower bands indicate that there are stronger CO-CO interactions in the adlayer, which are associated to a better ordered structure.⁴¹ The corresponding Stark tuning effects for linear adsorbed CO under this condition is $dv_{\text{CO}^l}/dE = 33 \text{ cm}^{-1} \text{ V}^{-1}$. Finally, for CO adsorbed in acid and analyzed in acidic solutions, the spectrum shows CO^l band at $\sim 2057 \text{ cm}^{-1}$ (with its respective $dv_{\text{CO}^l}/dE = 35 \text{ cm}^{-1} \text{ V}^{-1}$) and CO^b band $\sim 1828 \text{ cm}^{-1}$.⁷ The band frequencies and Stark effects are in perfect agreements with previous results from our laboratory in HClO_4 solution⁷ and those from Stamenkovic *et al.*⁴² and Zou *et al.*⁴³ for Pt(111)/CO in H_2SO_4 solution. On the other hand, Akemann *et al.*⁴⁴ report a Stark effect for CO on Pt(111) under HClO_4 solution about $20 \text{ cm}^{-1} \text{ V}^{-1}$ and $50 \text{ cm}^{-1} \text{ V}^{-1}$ for CO^l and CO^b , respectively. In these spectra, an additional band appears at $\sim 1256 \text{ cm}^{-1}$, which has been assigned to adsorbed sulfate in a three-fold coordinated symmetry site.⁴⁵ Similar results were obtained for the other two electrodes, which are summarized in Table 1.

From the FTIR results, it can be shown that the CO adlayer formed in acidic solution and analyzed in NaOH has maintained their structural characteristics when long-range (111) domains are present. For long-range (111) terraces, the Stark tuning effect for this adlayer examined in acid and alkaline conditions is the same, within the experimental error. Additionally, if the effect of the SHE potential is taken into account in the observed frequencies, it can be concluded that the band positions are in perfect agreement with those expected. For the Pt(17 17 15) electrode (as an example, Fig. 8 and Table 1), the expected change in frequency of the CO band from acid to alkaline solutions is 25 cm^{-1} . This value has been calculated using the Stark tuning slope of $\sim 34 \text{ cm}^{-1} \text{ V}^{-1}$ for CO^l recorded in acid solution and using the difference in SHE potential for the corresponding pH change (a pH change of 12 units, is equivalent to a change in SHE potential of ca. 0.70 V). The experimental difference is 24 cm^{-1} , which is, within the experimental error, equal to the predicted value. Moreover, an excellent linear fitting to the frequencies in acidic and alkaline solutions for this adlayer is obtained for the Pt(111) and Pt(17,17,15) electrodes (figure 9). This fact indicates that the identity of the adlayer is preserved upon transferring from the acid to the alkaline solution. Additionally, the higher frequencies for this adlayer form in acidic solutions as compared to that formed in alkaline solutions is associated with a layer with higher CO-CO interactions. Since the coverage is the same for both cases, the major difference should be then related to a lower number of defects present in the adlayer formed in acid solutions, in agreement with the proposed interpretation of the voltammetric results. For shorter (111) domains (as an example for Pt(443) in Fig. 8 and Table 1) although the frequency values are very similar, their Stark effect values suggested that short (111) domains contribute less to preserve the CO adlayer structure after transfer from the acid environment to the alkaline one.

4. Concluding remarks.

The results presented here indicate that the structure of the adlayer and the mobility of CO over the surface play a significant role in the kinetics of CO oxidation. Moreover, both CO mobility and CO adlayer surface structure depend on the SHE electrode potential, which can be seen as the main source of the observed differences. The structure of the adlayer plays a significant role in the onset of CO oxidation, due to the oxidation mechanism. CO oxidation through a Langmuir-Hinshelwood mechanism requires the presence of defects so that the oxidation can start.^{11, 35} The higher the number of defects are initially present in the adlayer, the lower the onset for oxidation is. Thus, the apparent higher catalytic activity of the Pt electrode for the oxidation of CO in alkaline media, which has been proposed because the lower onset for CO oxidation recorded in voltammetry as compared with acidic solutions, is a consequence of the a higher number of defects in the CO adlayer, and not a true catalytic effect. The lower number of defects (or the larger domains observed when CO is present in solution²³) can be also a consequence of the different mobility of CO. In acidic environments, the high mobility of the CO on the surface allows the coalescence of the different domains of the ordered adlayer formed in the initial stages of CO adsorption, giving rise to a well-ordered adlayer. In contrast, the low mobility of the CO does not allow such a process, producing smaller domains and a higher number of domain boundaries, linked to a lower onset for the CO stripping process.

Once the oxidation has started, the mobility of CO is which determines how this process occurs (i.e., only one or several peaks appear in the voltammetric profile). If the mobility of adsorbed CO is high as compared to the scan rate, CO will diffuse over the surface and eventually will find a site in which its oxidation is favorable. Thus, a single peak is always observed. On the other hand, several peaks are observed for the low mobility case because CO can be only oxidized on a particular site when the proper potential is reached. It should be recalled that, for low CO

coverages, whose oxidation is not affected by the presence of defects in the adlayer, CO oxidation in acidic solution takes place at lower potentials in alkaline media. For Pt(111) stepped surfaces, low CO coverages are oxidized at ca. 0.65-0.70 V vs RHE in acidic solutions,⁹ whereas similar coverages are only oxidized at ca. 0.8 V vs RHE in alkaline solutions.²⁰ This different behavior can be easily explained because of the different mobility or local effects in CO adlayer depending on both the electrode potential or charge on electrode surface.

Additionally, the constancy in the physical properties of the CO adlayer with the electrode potential, as those observed in the experiments where CO is adsorbed in acidic solutions and transferred to alkaline media, makes possible the comparison with the UHV values. This comparison can be used in the estimation of other physical properties of the electrode, as proposed by Weaver.⁴⁶⁻⁴⁷ In these works, the pzc of the CO covered Pt(111) electrode was assessed using the work function measured in UHV for the Pt(111) electrode covered with CO and D₂O⁴¹. According to the model proposed by Trasatti,⁴⁸ the relationship between the work function and the pzc (E_{pzc}) is:

$$E_{pzc} = \Phi / e - E_{Ref} \quad (1)$$

where Φ is the work function, e is the elementary charge and E_{Ref} the so-called “absolute” potential of the reference electrode. Using the measured value of the work function for the CO covered surface (5.6 eV) and an estimation of absolute potential of the reference electrode (4.6 V), a value of 1.0 V was obtained.⁴⁶ The value of 4.6 V for the absolute potential was selected because it was in the middle of the normal range of proposed values (between 4.4 and 4.8 V).⁴⁸ In the present case, we have found that the infrared properties of the CO adlayer are constant over a large potential window. Thus, these data can be used to obtain an estimation of the pzc of the CO covered electrode. The Stark effect shows that the frequencies for CO are very dependent on the electrode

potential. At the pzc of for the electrode, the frequencies for CO should be the same than those measured in UHV, since the absolute potential for both situations is the same. For the Pt(111) covered with CO and D₂O, the band frequency for linear CO is 2089 cm⁻¹.⁴¹ The linear regression obtained in figure 9 can be used to obtain such value. From the linear regression, the extrapolated value for the potential at which the band frequency for linear CO is 2089 cm⁻¹ is 0.94 V vs. SHE. A similar result is obtained with the Pt(17,17 15) surface which contains large (111) terraces. Our value is ca. 0.03 V lower than that proposed by Weaver.⁴⁶ However, it should be stressed that the latter value requires the knowledge of the exact value of the absolute reference potential, whereas the calculation presented here does not depend on that value. In fact, a value for E_{Ref} of 4.66 V would have given the same pzc. In conclusion, similar adlayer lead to the same interfacial properties.

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Table 1. Frequencies of adsorbed CO at 0.100 V and their corresponding Stark tuning slopes for adsorbed CO on different platinum single crystal electrodes in different conditions.

Electrodes	$\nu_{\text{CO}}^{\text{I}}/\text{cm}^{-1}$	$\nu_{\text{CO}}^{\text{b}}/\text{cm}^{-1}$	$d\nu_{\text{CO}}^{\text{I}}/dE/\text{cm}^{-1}\text{V}^{-1}$
Pt(111) ^a	2025	1795	+45
Pt(111) ^b	2029	1794	+33
Pt(111) ^c	2057	1828	+35
Pt(17 17 15) ^a	2018	1792	+47
Pt(17 17 15) ^b	2037	1830	+42
Pt(17 17 15) ^c	2061	1827	+34
Pt(443) ^a	2012	1827	+91
Pt(443) ^b	2033	1839	+49
Pt(443) ^c	2062	1831	+35

^aAdsorbed in 0.1 M NaOH examined in 0.1 M NaOH.

^bAdsorbed in 0.5 M H₂SO₄ and examined in 0.1 M NaOH.

^cAdsorbed in 0.1 M H₂SO₄ and examined in 0.1 M H₂SO₄.

Figures

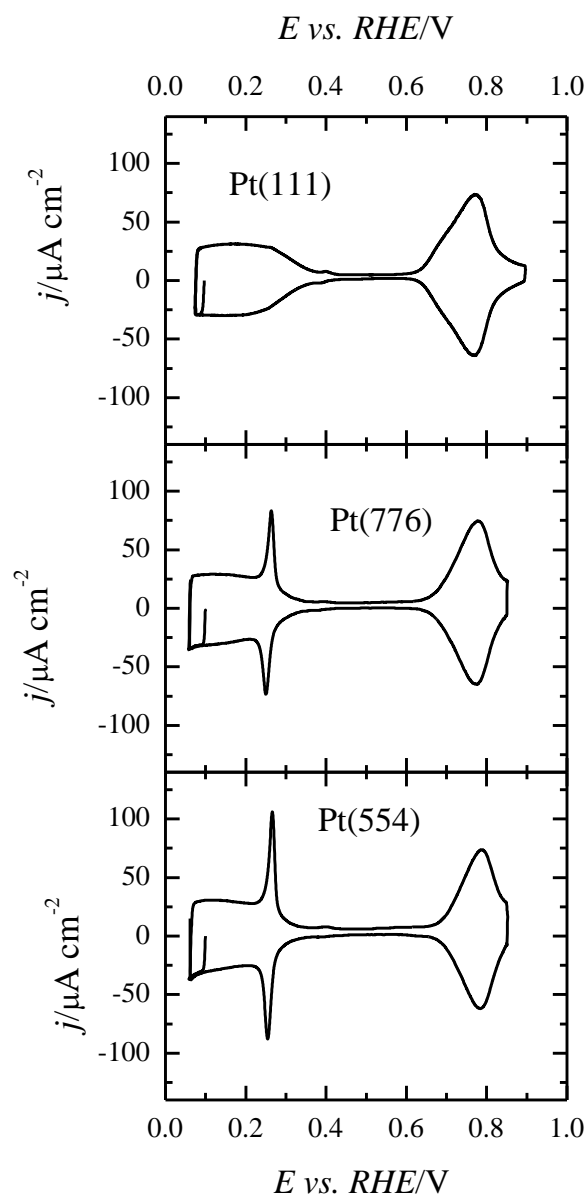


Figure 1. Voltammetric profiles for platinum single crystal electrodes in 0.1 M NaOH. Scan rate: 0.05 V s^{-1} .

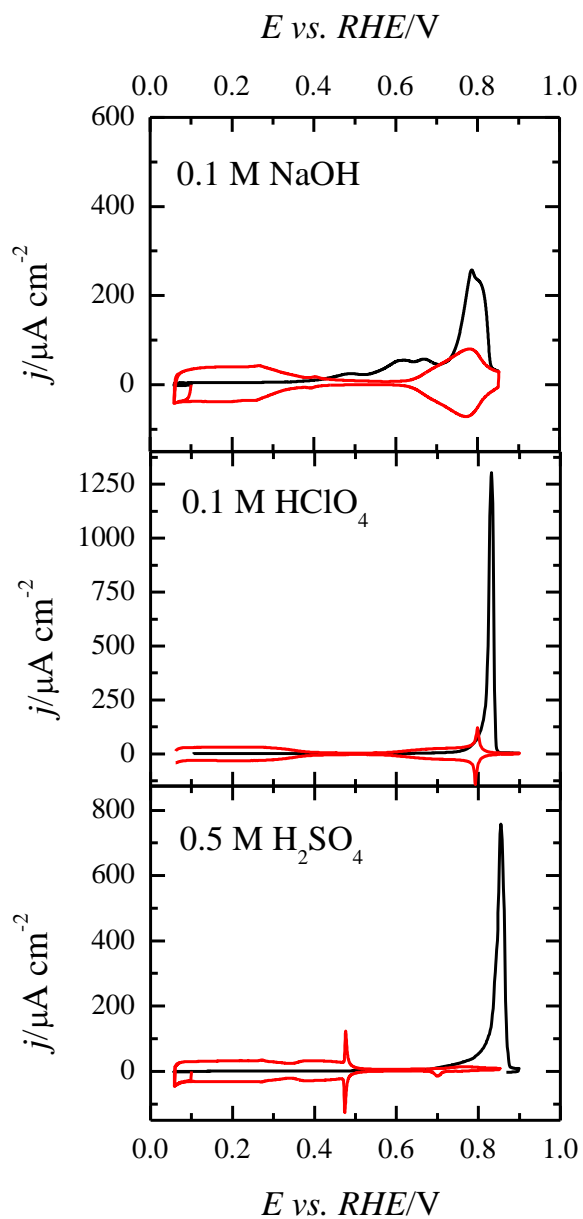


Figure 2. Voltammetric profiles for CO stripping (full line) and the clean surface after CO oxidation (dashed line) on a Pt(111) electrode in different electrolytic solutions. In all cases, CO has been adsorbed at 0.1 V in the same electrolytic solution. Scan rate: 0.05 V s^{-1}

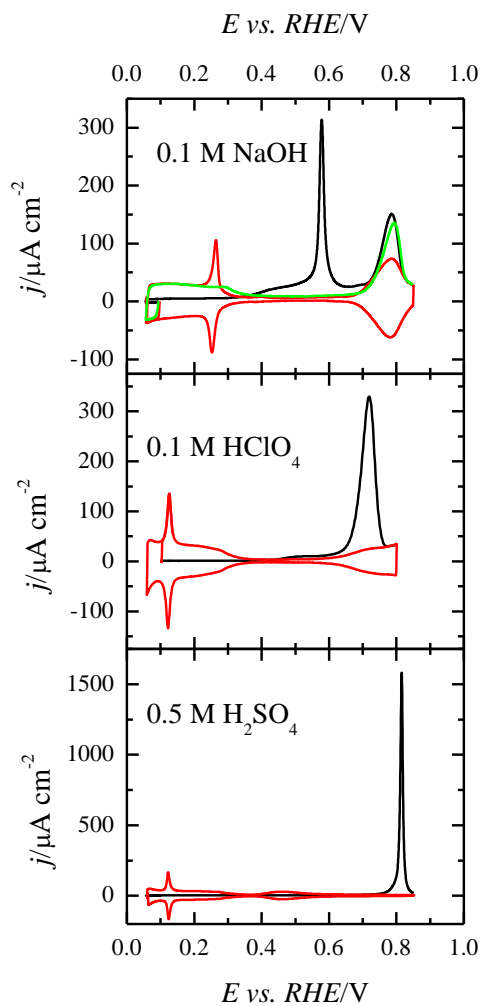


Figure 3. Voltammetric profiles for CO stripping (black line) and the clean surface after CO oxidation (red line) on a Pt(554) electrode in different electrolytic solutions. In all cases, CO has been adsorbed at 0.1 V in the same electrolytic solution. The green line in the top panel represents the voltammetric profile for CO stripping when CO is only adsorbed on the step sites. Scan rate: 0.05 V s^{-1} .

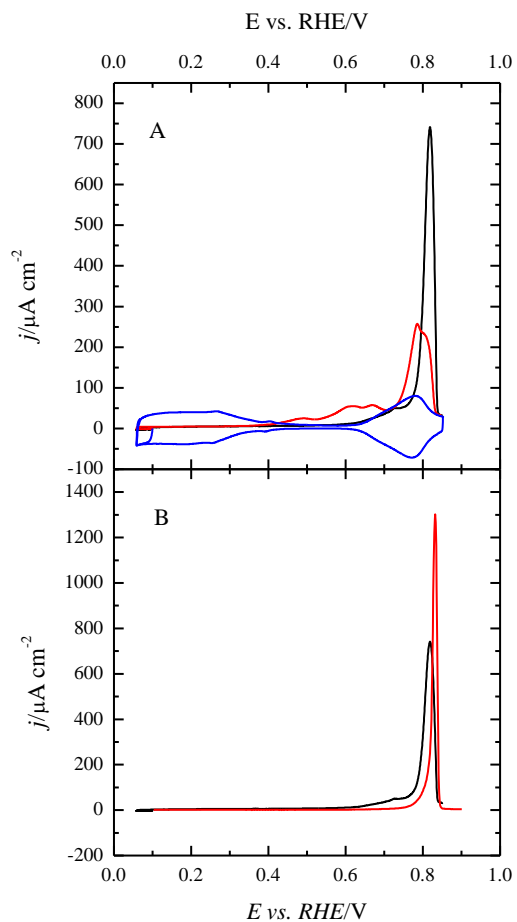


Figure 4. A) Voltammetric profile on a Pt(111) electrode for CO stripping in 0.1 M NaOH for CO adsorbed at 0.1 V in 0.5 M H_2SO_4 (black line) and for CO adsorbed at 0.1 V in 0.1 M NaOH (red line). The blue line shows the profile of the Pt(111) electrode after CO stripping. B) Voltammetric profile on a Pt(111) electrode for CO stripping in 0.1 M NaOH for CO adsorbed at 0.1 V in 0.5 M H_2SO_4 (black line) and for CO stripping in 0.1 M HClO_4 for CO adsorbed at 0.1 V in 0.1 M HClO_4 (red line). Scan rate: 0.05 V s^{-1} .

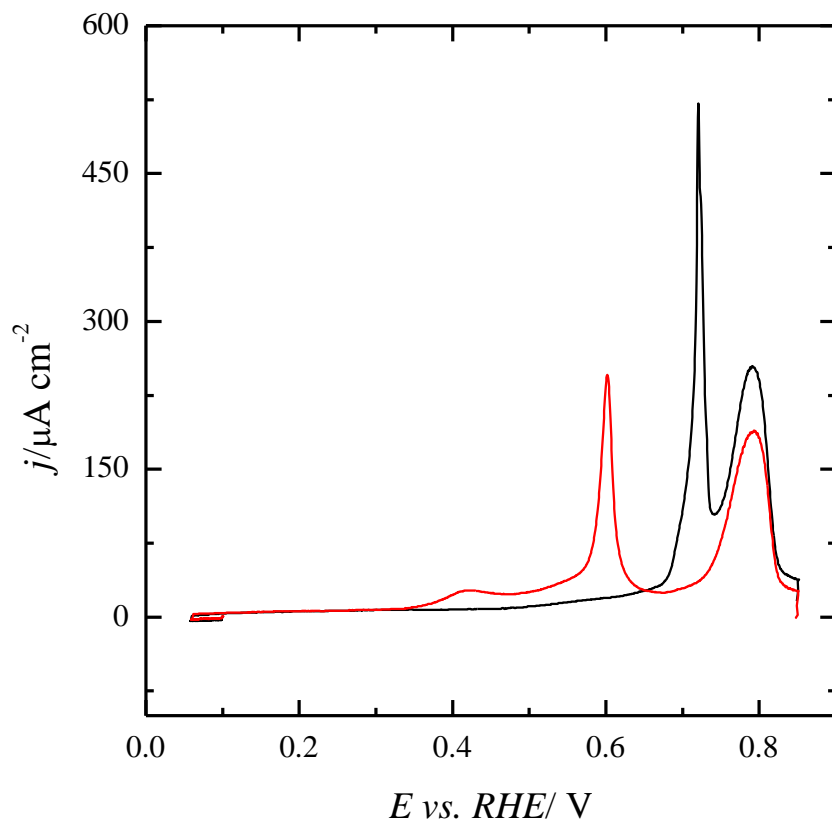


Figure 5. Voltammetric profile on a Pt(776) electrode for CO stripping in 0.1 M NaOH for CO adsorbed at 0.1 V in 0.5 M H₂SO₄ (black line) and for CO adsorbed at 0.1 V in 0.1 M NaOH (red line). Scan rate: 0.05 V s⁻¹.

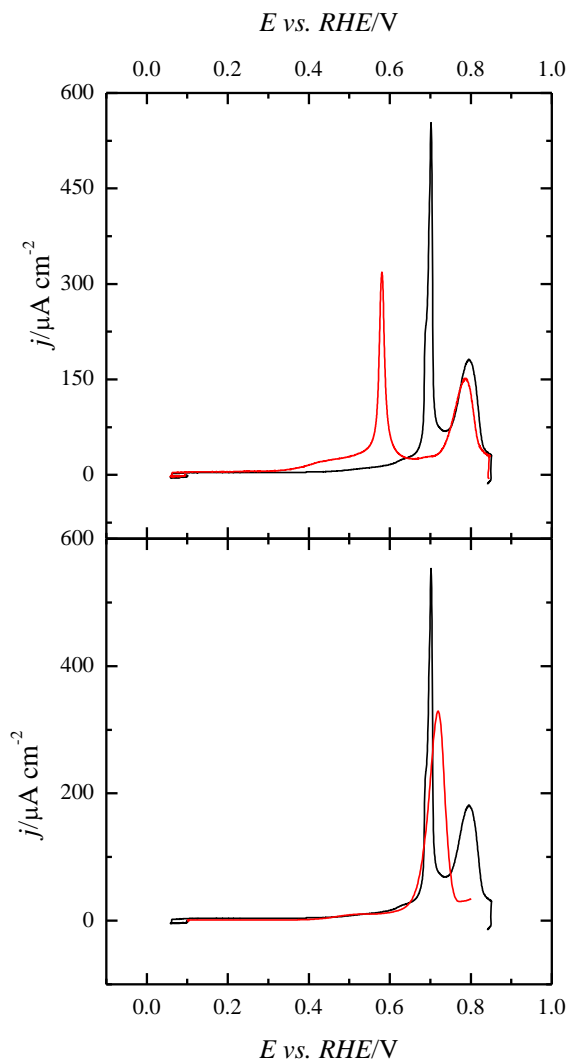


Figure 6. A) Voltammetric profile on a Pt(554) electrode for CO stripping in 0.1 M NaOH for CO adsorbed at 0.1 V in 0.5 M H₂SO₄ (black line) and at 0.1 V in 0.1 M NaOH (red line). B) Voltammetric profile on a Pt(554) electrode for CO stripping in 0.1 M NaOH for CO adsorbed at 0.1 V in 0.5 M H₂SO₄ (black line) and for CO stripping in 0.1 M HClO₄ for CO adsorbed at 0.1 V in 0.1 M HClO₄ (red line). Scan rate: 0.05 V s⁻¹.

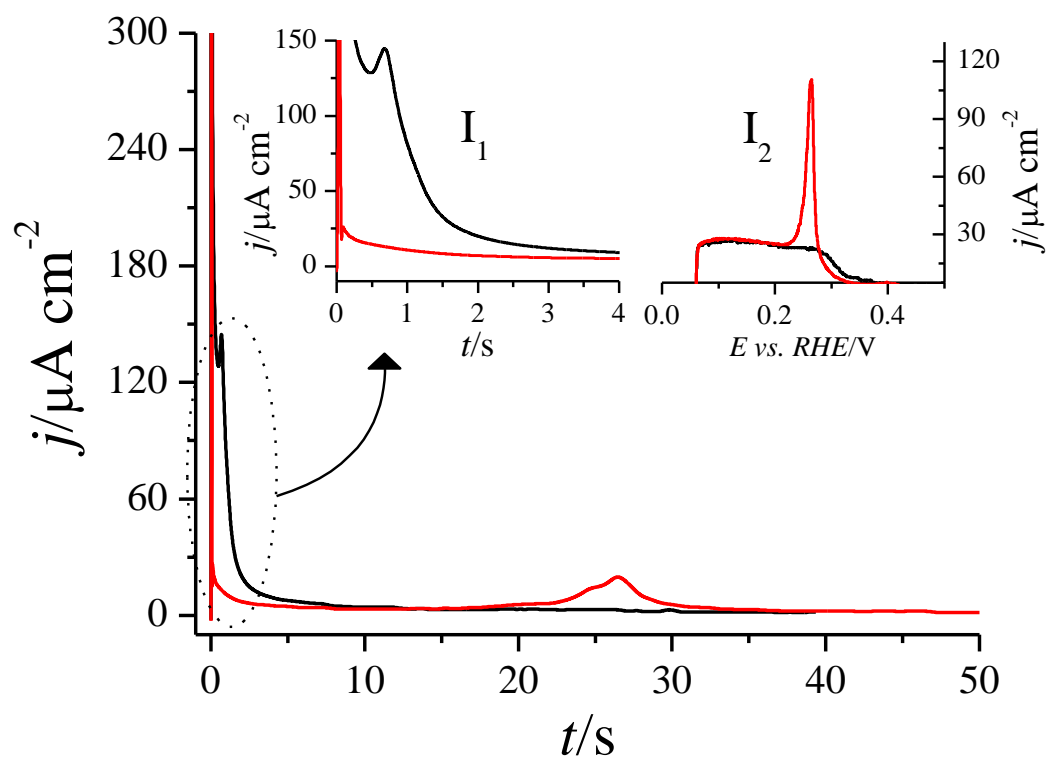


Figure 7. Current transient for CO stripping at 0.58 V on a Pt(554) electrode in 0.1 M NaOH for CO dosed at 0.1 V in 0.1 M NaOH (black line) and for CO dosed at 0.1 V in 0.5 M H₂SO₄ (red line). Insert: I₁) Magnification of the initial part of the transient. I₂) Voltammetric profile of the Pt(554) electrode after the transients (black line) and the clean surface (red line).

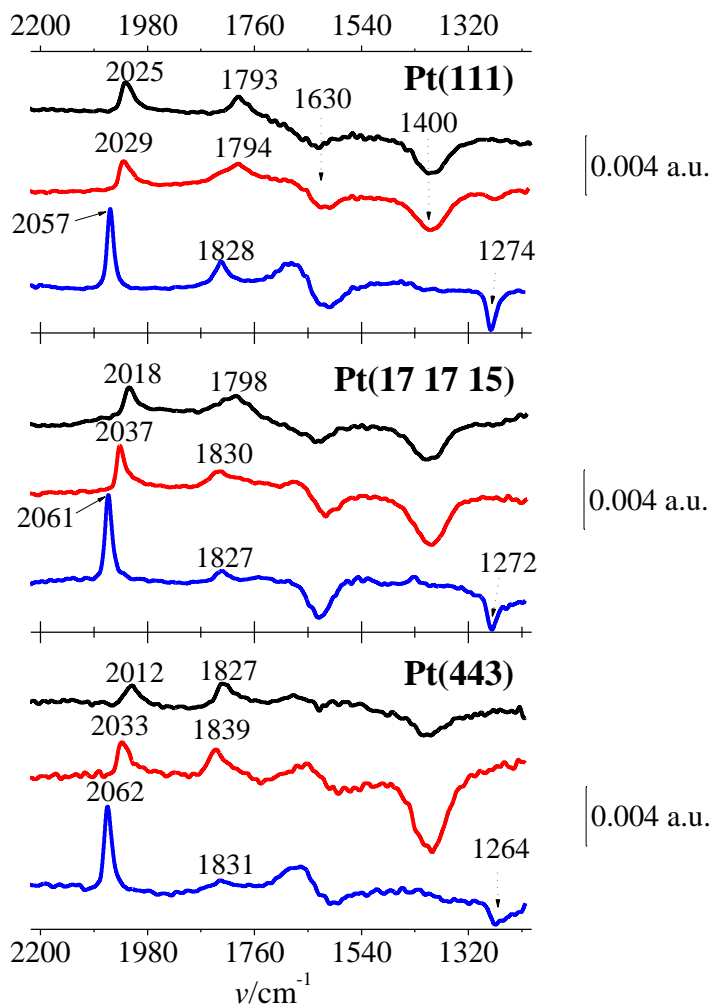


Figure 8. *In situ* FTIR spectra recorded for full CO monolayer for different Pt electrodes. Black line: CO adsorbed at 0.1 V in 0.1 M NaOH and spectra acquired at 0.1 V in NaOH. Red line: CO adsorbed at 0.1 V in 0.5 M H₂SO₄ and spectra acquired at 0.1 V in NaOH. Blue line: CO adsorbed at 0.1 V in 0.1 M H₂SO₄ and spectra acquired at 0.1 V in 0.1 M H₂SO₄. The reference spectra was always measured at 0.90 V in the respective solution.

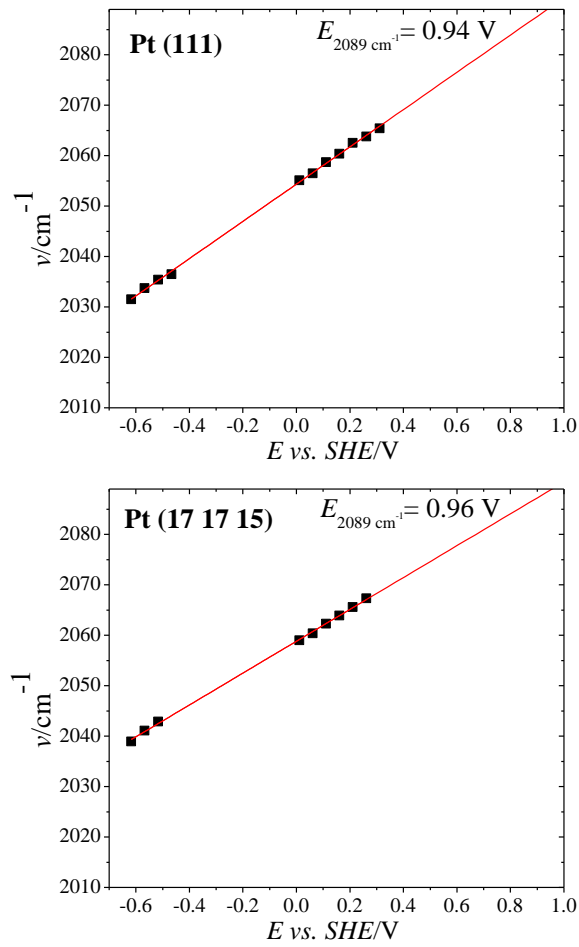


Figure 9. Frequencies for linear adsorbed CO vs. electrode potential (vs. SHE) for a CO layer formed at 0.1 V in H_2SO_4 solution studied both in acidic and alkaline solutions. The linear fits are also included, indicating the potential at which the extrapolated frequency would be 2089 cm^{-1} .

TOC

